

The First Total Synthesis of (\pm)-Cyclophostin and (\pm)-Cyclipostin P: Inhibitors of the Serine Hydrolases Acetyl Cholinesterase and Hormone Sensitive Lipase

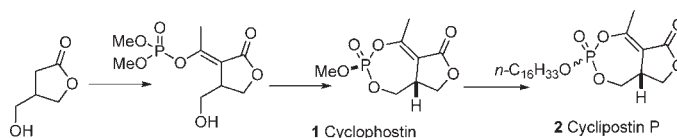
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ABSTRACT



Cyclophostin, a structurally unique and potent naturally occurring acetyl cholinesterase (AChE) inhibitor, and its unnatural diastereomer were prepared in 6 steps and 15% overall yield from hydroxymethyl butyrolactone. The unnatural diastereomer of cyclophostin was converted into cyclipostin P, a potent naturally occurring hormone sensitive lipase (HSL) inhibitor, using a one pot dealkylation–alkylation process. The inhibition [IC_{50}] of human AChE by cyclophostin and its diastereomer are reported, as well as constituent binding (K_i) and reactivity (k_2) constants.

Cyclophostin **1** (Figure 1) is a novel bicyclic organophosphate isolated from a fermentation solution of *Streptomyces lavendulae* (strain NK901093).¹ The natural product **1** showed potent inhibition of acetyl cholinesterase (AChE) from the housefly (CSMA strain) and the brown plant hopper with a reported IC_{50} of 7.6×10^{-10} M. The structure of cyclophostin was first assigned by spectroscopic methods and then confirmed by single crystal X-ray diffraction studies as a bicyclic structure with a seven-membered cyclic enol-phosphate triester fused to a butyrolactone ring. There are chirality centers at both C3a and the phosphorus atom (6). The absolute configurations were determined to be 3a*R*,6*S*.²

The unusual bicyclic enolphosphate is also found in the family of structurally related natural products, named the

cyclipostins **2**.³ The cyclipostins **2** possess a core structure similar to that of cyclophostin, but differ in the phosphate ester. The cyclipostins **2** are phosphate esters of long chain lipophilic alcohols of various lengths and structures, and all are potent inhibitors of hormone sensitive lipase (HSL). In addition, cyclophostin and the cyclipostins are probably biosynthetically related to A-factor **3** and the virginia butanolides,⁴ and the PLA₂ inhibitor **4**.⁵

The significance of cyclophostin and the cyclipostins is due in part to their unique structures and the enzymes they

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(2) The configuration for cyclophostin named in ref 1, 3a*R*,6*S*, is actually incorrect. A common error is to treat P^+-O^- as $P=O$ and hence assign the priority incorrectly. Natural cyclophostin is actually 3a*R*,6*R*. See: Cahn, R. S.; Ingold, C.; Prelog, V. *Angew. Chem., Int. Ed.* **1966**, *5*, 385. The configuration is correctly named in chemical abstracts.

(3) (a) Wink, J.; Schmidt, F.-R.; Seibert, G.; Aretz, W. *J. Antibiot.* **2002**, *55*, 472. (b) Vértesy, L.; Beck, B.; Brönstrup, M.; Ehrlich, K.; Kurz, M.; Müller, G.; Schummer, D.; Seibert, G. *J. Antibiot.* **2002**, *55*, 480.

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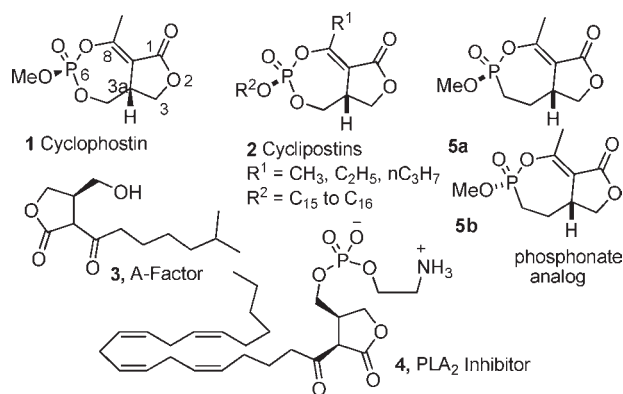


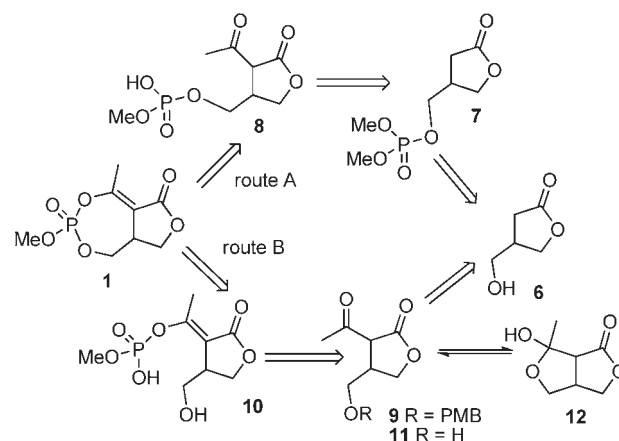
Figure 1. Cyclophostin and related structures.

inhibit. Interest in AChE has re-emerged due its role as a therapeutic target for Alzheimer's disease,⁶ myasthenia gravis,⁷ and glaucoma,⁸ whereas HSL is a therapeutic target for type II diabetes.⁹

We recently reported the synthesis of a phosphonate analog **5a** of cyclophostin.¹⁰ The activity of the phosphonates was ≥ 100 -fold less than the value reported for cyclophostin. The *trans* isomer (H and OMe) **5b** was 10-fold more active (IC_{50} of $3 \mu M$ human AChE) than the *cis* isomer **5a** (IC_{50} of $30 \mu M$ human AChE). Since the natural product has the *cis* (H and OMe) configuration, the unnatural isomer may well prove more potent. To accurately compare the activities of cyclophostin, the phosphonate analog, and their diastereomers with a detailed kinetic analysis, we needed reasonable quantities of the natural product. Furthermore, we proposed that cyclophostin would be an excellent precursor for the synthesis of the family of cyclipostins. Herein, we report the first synthesis of (\pm) cyclophostin and conversion into (\pm) cyclipostin P.

A retrosynthetic analysis (Scheme 1) of the bicyclic phosphate **1** suggested that the cyclic enolphosphate could be formed by either condensation of the acetyl group (as the enol) with a phosphoric acid via intermediate **8**

Scheme 1. Cyclophostin Retrosynthetic Analysis



(route A) or conversely condensation of the primary alcohol with an enolphosphoric acid via intermediate **10** (route B). Both intermediates can be formed by C-acylation of derivatives of hydroxymethyl lactone **6**. The lactone **6** and various derivatives are available in the racemic modification¹¹ and either enantiomer.¹² It was thought necessary to protect the hydroxyl of lactone **6** prior to C-acylation to avoid complications arising from cyclization of the acetyl lactone **11** to the hemiketal **12**.^{12b} The most expedient route would be to introduce the phosphate early in the synthesis, simultaneously protecting the hydroxyl group.

The racemic hydroxy lactone **6** was prepared using published methods.¹¹ The hydroxyl was phosphorylated using dimethyl bromophosphate, prepared *in situ* by reaction of trimethyl phosphite with bromine, to give the phosphate **7** (Scheme 2). The phosphorylated butyrolactone **7** was deprotonated with 1 equiv of LiHMDS in THF, and the resulting enolate was acylated with acetyl chloride.^{4a,5} Initially, mixtures of the acetyl lactone **14** and the enolacetate **13** were observed, so an excess of acetyl chloride was added to ensure the complete acylation giving enolacetate **13** in 65% yield as a mixture of two geometrical isomers. The geometrical isomers could be separated, but were generally carried through to the next step as a mixture. Deacetylation of enolacetates **13** was achieved using a catalytic amount of DMAP in MeOH to give the acetyl lactone **14** in 62% yield.

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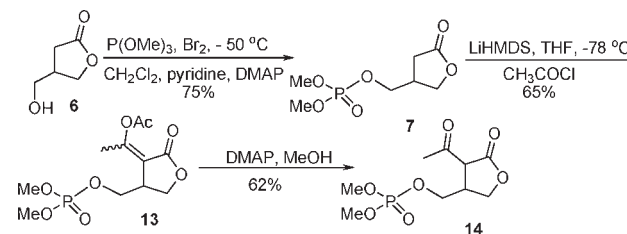
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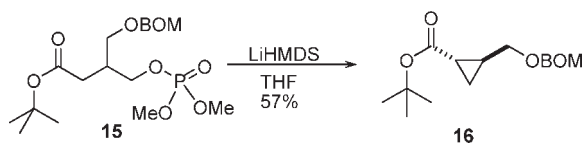
(11) For examples, see: (a) Gadir, S. A.; Smith, Y.; Taha, A. A.; Thaller, V. *J. Chem. Res. (S)* **1986**, 222. (b) Yoda, H.; Mizutani, M.; Takabe, K. *Synlett* **1998**, 855. (c) Sengoku, T.; Suzuki, T.; Kakimoto, T.; Takahashi, M.; Yoda, H. *Tetrahedron* **2009**, *65*, 2415.

(12) For examples, see: (a) Banfi, L.; Basso, A.; Guanti, G.; Zannetti, M. T. *Tetrahedron: Asymmetry* **1997**, *8*, 4079. (b) Parsons, P. J.; Lacrouts, P.; Buss, A. D. *Chem. Commun.* **1995**, 437. (c) Crawforth, J. M.; Fawcett, J.; Rawlings, B. *J. Chem. Soc., Perkin Trans. 1* **1998**, 1721. (d) Posner, G. H.; Wietzberg, M.; Jew, S.-S. *Synth. Commun.* **1987**, *17*, 611. (e) Takabe, K.; Mase, N.; Matsumura, H.; Hasegawa, T.; Iida, Y.; Kuribayashi, H.; Adachi, K.; Yoda, H.; Ao, M. *Bioorg. Med. Chem. Lett.* **2002**, *12*, 2295. (f) Mori, K.; Chiba, N. *Liebigs Ann. Chem.* **1989**, 957. (g) Takabe, K.; Tanaka, M.; Sugimoto, M.; Yamada, T.; Yoda, H. *Tetrahedron: Asymmetry* **1992**, *3*, 1385.

Scheme 2. Synthesis of Primary Phosphate

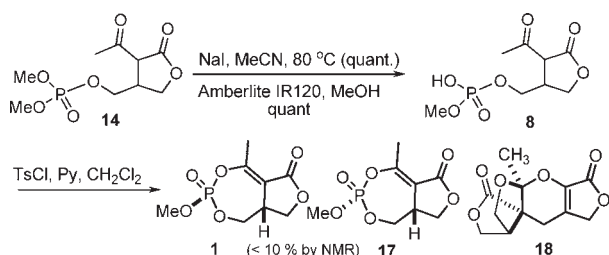


Scheme 3. Formation of Cyclopropanes



The successful C-acylation of the lactone **7** was both gratifying and somewhat surprising. It is well-known that the enolates derived from γ -phosphoryloxy carboxylates cyclize to form cyclopropanes.¹³ Indeed, when the acyclic γ -phosphoryloxy carboxylate **15** was treated with LiHMDS and acetyl chloride (Scheme 3), the only isolable product was the cyclopropane **16**. In contrast, a solution of the enolate of butyrolactone **7** was stable in the absence of an external electrophile up to $-20\text{ }^\circ\text{C}$ for 1 h. At higher temperatures decomposition to intractable products was observed.

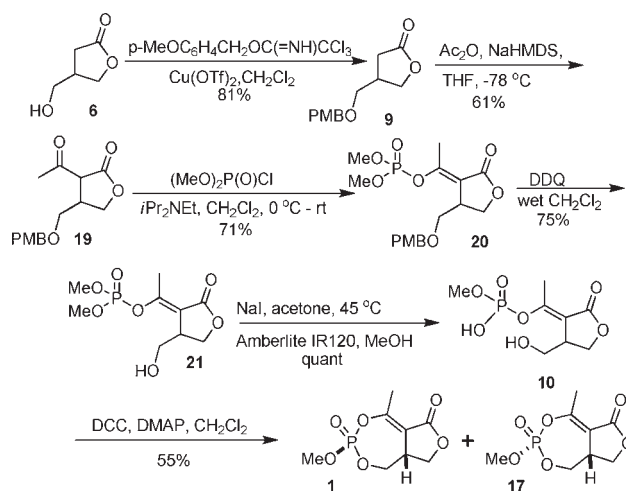
Scheme 4. Attempted Cyclization of the Primary Phosphate



With the desired 2-acetyl butyrolactone intermediate **14** in hand, the final step of the cyclophostin synthesis was explored (Scheme 4). The intermediate **14** was demethylated using 1 equiv of sodium iodide in refluxing acetonitrile solution to give the corresponding sodium salt in quantitative yield.¹⁴ The sodium salt was protonated using Amberlite (sulfonic acid) resin to yield the phosphoric acid **8**. Attempted cyclization via intramolecular condensation of the monophosphoric acid moiety with the enol form of the acetyl group using EDC, HOBt, and Hunig's base and related reactions were not successful.¹⁵ These conditions had been used previously in the synthesis of the phosphonate analog.¹⁰ Other reaction conditions were examined. Reaction with TsCl and pyridine gave a crude product with signals at -7.6 , -8.0 , and -11.7 ppm in the ^{31}P NMR spectrum (among others). After aqueous

workup, a significant amount of mass was lost in the organic soluble crude product. The peak at -8.0 had disappeared, and only the peaks at -7.6 and -11.7 ppm remained in the crude product. Chromatographic separation of the product mixture using preparative TLC yielded cyclophostin **1** and its unnatural diastereomer **17** in very low yield ($< 10\%$ combined yield). The identity of cyclophostin was confirmed by comparing the ^1H and ^{31}P NMR data with those reported for the natural product.¹ A crystalline tetracyclic byproduct **18** was also isolated, and the structure was confirmed by an X-ray diffraction study.¹⁶

Scheme 5. Synthesis and Cyclization of the Enolphosphate



The lack of success in finding reaction conditions for the cyclization of **8** led us to examine an alternate route **B** (Scheme 5). The hydroxyl of lactone **6** was protected as a *p*-methoxybenzyl (pmb) ether **9** by the copper(II) triflate catalyzed reaction with *p*-methoxybenzyloxy trichloroacetimidate.¹⁷ Deprotonation with NaHMDS and acylation with acetic anhydride furnished the acetyl lactone **19** directly. Selective phosphorylation of 2-acetyl butyrolactone **19** was accomplished by reaction with dimethyl chlorophosphate using a procedure reported to give the *E*-enolphosphate.^{18,19}

The pmb ether was removed using DDQ in wet CH_2Cl_2 , and the enolphosphate **21** was selectively monodemethylated using 1 equiv of sodium iodide in acetone at $45\text{ }^\circ\text{C}$. The sodium salt was protonated with Amberlite IR 120 resin to generate the corresponding phosphoric acid **10**. Reaction of the phosphoric acid **10** with DCC and DMAP in CH_2Cl_2 successfully formed the cyclic enolphosphates **1**

(13) For examples, see: (a) Jacks, T. E.; Nibbe, H.; Wiemer, D. F. *J. Org. Chem.* **1993**, *58*, 4584. (b) Singh, A. K.; Rao, M. N.; Simpson, J. H.; Li, W.-S.; Thornton, J. E.; Kuehner, D. E.; Kacsur, D. J. *Org. Process Res. Dev.* **2002**, *6*, 618. (c) Krawczyk, H.; Wasek, K.; Kedzia, J.; Wojciechowski, J.; Wolf, W. M. *Org. Biomol. Chem.* **2008**, *6*, 308.

(14) The reaction of phosphates with NaI is generally performed in acetone, but other solvent including MeCN also work well. See: Zervas, L.; Dilaris, I. *J. Am. Chem. Soc.* **1955**, *77*, 5354.

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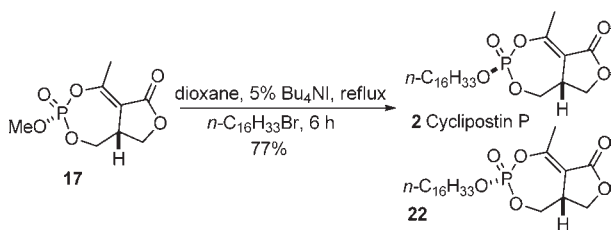
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and **17** as a 1:1 mixture. The diastereomers were separated using silica gel chromatography to give natural cyclophostin **1** and its diastereomer **17** in 55% combined yield.

Cyclophostin and its diastereomer are potential precursors for the synthesis of the family of cyclipostins by ester exchange. We opted to explore a novel one pot process for this conversion (Scheme 6). The unnatural diastereomer **17** was treated with hexadecyl bromide (10 equiv) and catalytic tetrabutylammonium iodide (TBAI) in refluxing dioxane to give cyclipostin P **2** and its diastereomer **22** with 95% conversion and a 1:1 ratio. The diastereomers were separated by column chromatography to give a 77% combined yield. The reaction of either **2** or **22** with hexadecyl bromide and catalytic TBAI in refluxing dioxane resulted in a mixture of both **2** and **22** in a 1:1 ratio. There was no reaction in the absence of TBAI.

Scheme 6. Conversion of Cyclophostin into Cyclipostin P



It is proposed that the iodide cleaves the methyl (or alkyl) phosphate ester bond to give a phosphate anion which is then alkylated by the hexadecyl bromide. The formation of the anion would account for the scrambling of the stereochemistry at phosphorus.

(±)-Cyclophostin and its diastereomer were examined for inhibitory activity against human AChE using a modified Ellman assay.²⁰ Compounds **1** and **17** are potent inhibitors of human AChE: both have IC₅₀'s of ~40 nM (Table 1). This is only slightly weaker than the previously reported values for the natural product with AChE isolated from the insect.¹ This could be a reflection of either the difference in AChE source species or the selectivity of the single (natural) enantiomer versus the racemic (synthetic) mixture. There are two additional comparisons of interest. One is that although they are diastereomers, the IC₅₀'s for **1** and **17** are very similar. The other is that these IC₅₀'s represent a 10³–10⁴ improvement in potency relative to the corresponding (±)-phosphonate analogs **5**.^{10,18}

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Table 1. Inhibition of AChE by Cyclophostin and Related Compounds



no.	IC ₅₀	K _I	k ₂ , min ⁻¹	k ₋₂ , min ⁻¹
5a	30 μM	140 ± 60 μM	0.4 ± 0.1	4e-3
5b	3 μM	24 ± 6.0 μM	0.3 ± 0.06	0.02
1	45 nM	140 ± 72 nM	0.7 ± 0.02	0
17	40 nM	210 ± 140 nM	0.9 ± 0.5	0

To dissect this behavior, residual enzyme activities were obtained as a function of inhibitor concentration and incubation time for **1** and **17** and analyzed using the equation shown in Table 1 as described previously.¹⁸ Interestingly, the chemical (rate) constants k₂ and k₋₂ vary little among the four compounds, but the K_I, which represents the noncovalent binding event, reflects 100–1000-fold stronger binding by the phosphates than the phosphonates. Thus while chemical intuition would suggest that the carbon replacement of the endocyclic oxygen would cause a difference in reactivity, these data indicate that this substitution instead has profoundly adverse effects on binding. We speculate that this is due to a combination of additional hydrogen bonds (to the O) and subtle conformational differences in the bicyclic ring system. This observation points to the exquisite ability of enzymes to discriminate compounds in their active sites.

In conclusion, cyclophostin and its diastereomer have been prepared in six steps and 15% overall yield from lactone **6**. The diastereomer of cyclophostin was converted into cyclipostin P using a one pot process. Syntheses of additional examples of the cyclipostins and corresponding biological studies are ongoing and will be reported in due course.

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Supporting Information Available. Detailed experimental procedures and full spectroscopic data for all new compounds. The material is available free of charge via the Internet at <http://pubs.acs.org>.